TEST NAME: **PSc. 2.1 The Periodic Table Spring 2018**
TEST ID: **2239141**
GRADE: **09 - Ninth Grade - 12 - Twelfth Grade**
SUBJECT: **Life and Physical Sciences**
TEST CATEGORY: **School Assessment**

Student: _____

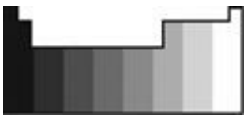
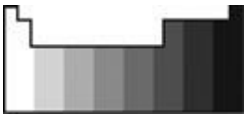
Class: _____

Date: _____

1. **Two elements in the same period on the Periodic Table of the Elements are MOST similar in their**

   A. atomic mass.
   B. number of electron shells.
   C. atomic size.
   D. valence electrons.

2. **The diagrams show part of the Periodic Table of the Elements. The shaded areas represent a general trend in the number of valence electrons in the elements. The lighter shading indicates fewer valence electrons and the darker shading indicates more valence electrons. Which diagram BEST shows the relationship between the elements and the number of valence electrons?**

   A. 
   B. 
   C. 
   D. 

3. **On the periodic table, helium (He) is placed at the top of group 18 (8A). What is the reason for helium being at this location on the periodic table?**

   A. Helium has a filled valence shell.
   B. Helium is a very reactive gas.
   C. Helium loses its electrons easily.
   D. Helium is a transition metal.

4. **The elements fluorine (F), chlorine (Cl), and iodine (I) are all in the same column on the periodic table. Which is the BEST explanation for this arrangement?**

A. They are all gases at room temperature.

B. They were all discovered in ancient times.

C. They have the same number of energy levels.

D. They have the same number of valence electrons.

5. Which of the metals listed below would be MOST reactive?

A. Cesium (Cs)

B. Sodium (Na)

C. Francium (Fr)

D. Potassium (K)