TEST NAME: **PSc 4.3 Covalent Bonding Spring 2018**
TEST ID: **2251632**
GRADE: **09 - Ninth Grade - 12 - Twelfth Grade**
SUBJECT: **Life and Physical Sciences**
TEST CATEGORY: **School Assessment**

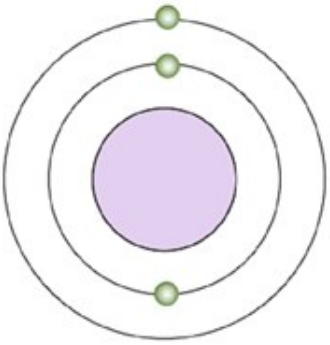
Student:

Class:

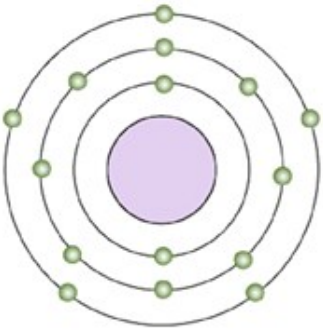
Date:

1. Which diagram represents an element that is likely to form **covalent** bonds?

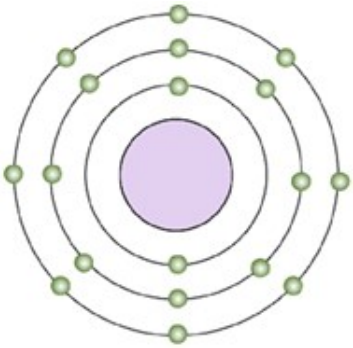
A.



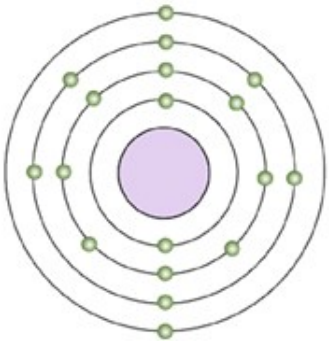
B.



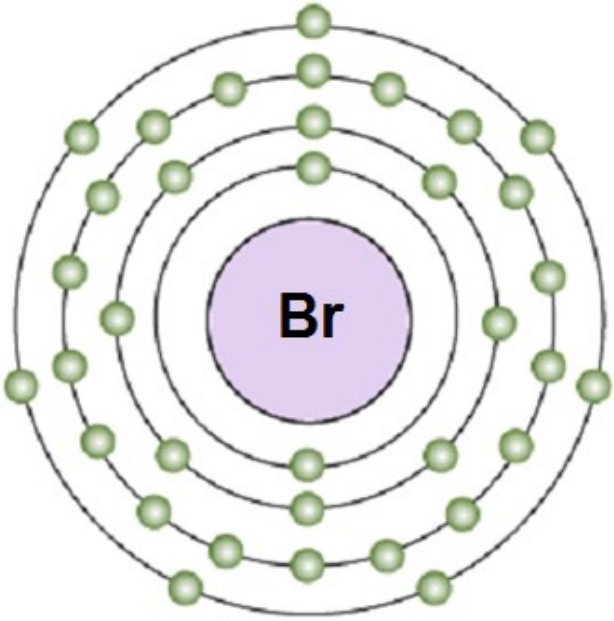
C.

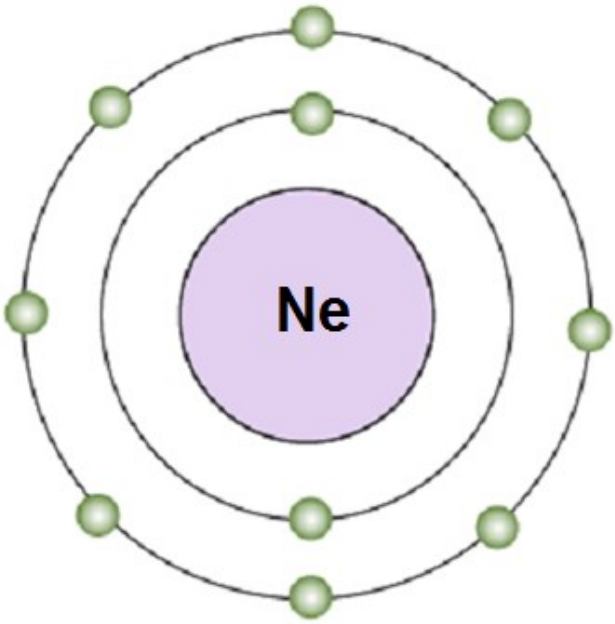


D.



2. **What best explains whether bromine (Br) or neon (Ne) is more likely to form a covalent bond?**





A. Bromine forms covalent bonds because it has **seven valence electrons**, but neon has **eight valence electrons** and already fulfills the octet rule (stable atom).

B. Bromine forms covalent bonds because it has many **electron shells,** but neon has only **two electron shells** and is tightly bound to its electrons.

C. Neon forms covalent bonds because it can share its **valence electrons,** but bromine has **seven valence electrons** and can **gain** only one more electron.

D. Neon forms covalent bonds because it has only **two electron shells,** but bromine has **many electron shells** and will lose electrons in order to fulfill the octet rule.

3. What is the formula of the compound Diphosphorus decasulfide?

A. P_2S_{10}

B. $2P_{10}S$

C. S_2P_{10}

D. $2S_{10}P$

4. What is the name of C_4N_5?

A. tetratcarbide pentanitride

B. tetracarbon pentanitride

C. carbon tetra nitride

D. Tetracarbon nitrate

5. What is the name of CO_3?

A. carbide trioxide

B. carbon trioxide

C. monocarbide trioxide

D. trioxide moncarbide